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Crash Course Chemistry #37: The History of Atomic Chemistry

Name: _____ Date: _____

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1. The Greek word atomos means _____ or indivisible.
2. Early Greek philosophers believed that different materials were composed of different types of _____.
3. The law of _____ of mass states that matter's mass remains unchanged even if it changes form.
4. Discharge tubes emit light when an _____ current passes through them.
5. Cathode rays were shown to carry a _____ charge.
6. J.J. Thompson called the negatively charged particles "corpuscles," which we now call _____.
7. Thompson's model of the atom is commonly referred to as the _____ pudding model.
8. Rutherford's gold foil experiment revealed that most of the atom is _____ space.
9. Rutherford concluded that the atom's positive charge is concentrated in a small area called the _____.
10. The positively charged fundamental particle found in the nucleus is called the _____.
11. Bohr's model of the atom is sometimes called the _____ model.
12. Each electron orbit in Bohr's model correlates with specific energy levels and _____.
13. Heisenberg introduced the idea that it is impossible to know both the momentum and _____ of an electron with certainty.
14. The quantum model of the atom describes the probable location of electrons using _____.
15. Because of the probabilistic nature of electron locations, the modern model of the atom is often called the _____ model.

Crash Course Chemistry #37: The History of Atomic Chemistry KEY

1. The Greek word atomos means **uncuttable** or indivisible.
2. Early Greek philosophers believed that different materials were composed of different types of **atoms**.
3. The law of **conservation** of mass states that matter's mass remains unchanged even if it changes form.
4. Discharge tubes emit light when an **electrical** current passes through them.
5. Cathode rays were shown to carry a **negative** charge.
6. J.J. Thompson called the negatively charged particles "corpuscles," which we now call **electrons**.
7. Thompson's model of the atom is commonly referred to as the **plum** pudding model.
8. Rutherford's gold foil experiment revealed that most of the atom is **empty** space.
9. Rutherford concluded that the atom's positive charge is concentrated in a small area called the **nucleus**.
10. The positively charged fundamental particle found in the nucleus is called the **proton**.
11. Bohr's model of the atom is sometimes called the **planetary** model.
12. Each electron orbit in Bohr's model correlates with specific energy levels and **orbitals**.
13. Heisenberg introduced the idea that it is impossible to know both the momentum and **position** of an electron with certainty.
14. The quantum model of the atom describes the probable location of electrons using **orbitals**.
15. Because of the probabilistic nature of electron locations, the modern model of the atom is often called the **cloud** model.

15 Questions

Answer Key Included

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