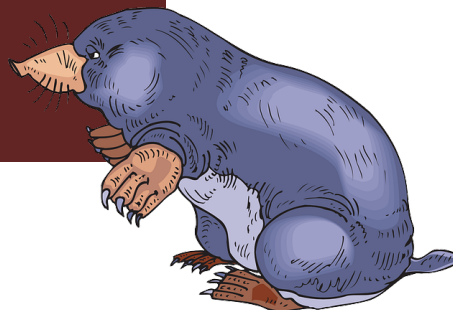
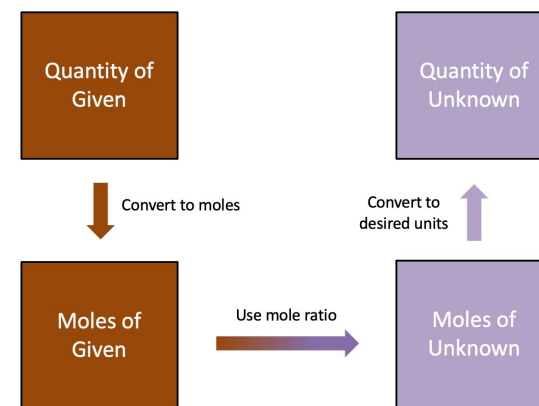
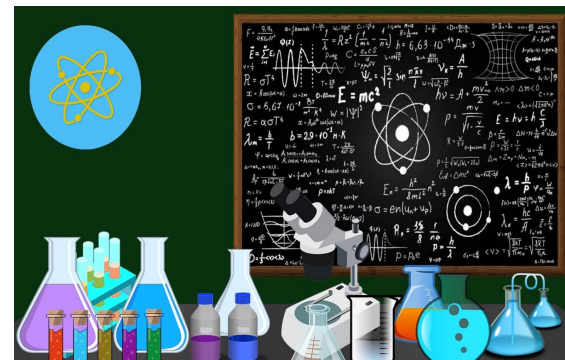


Stoichiometry

- What is Stoichiometry?
- Chemical Reactions and the Mole
- Steps to Stoichiometry
- Mole Ratio
- Fully Worked Practice Questions



Preview File

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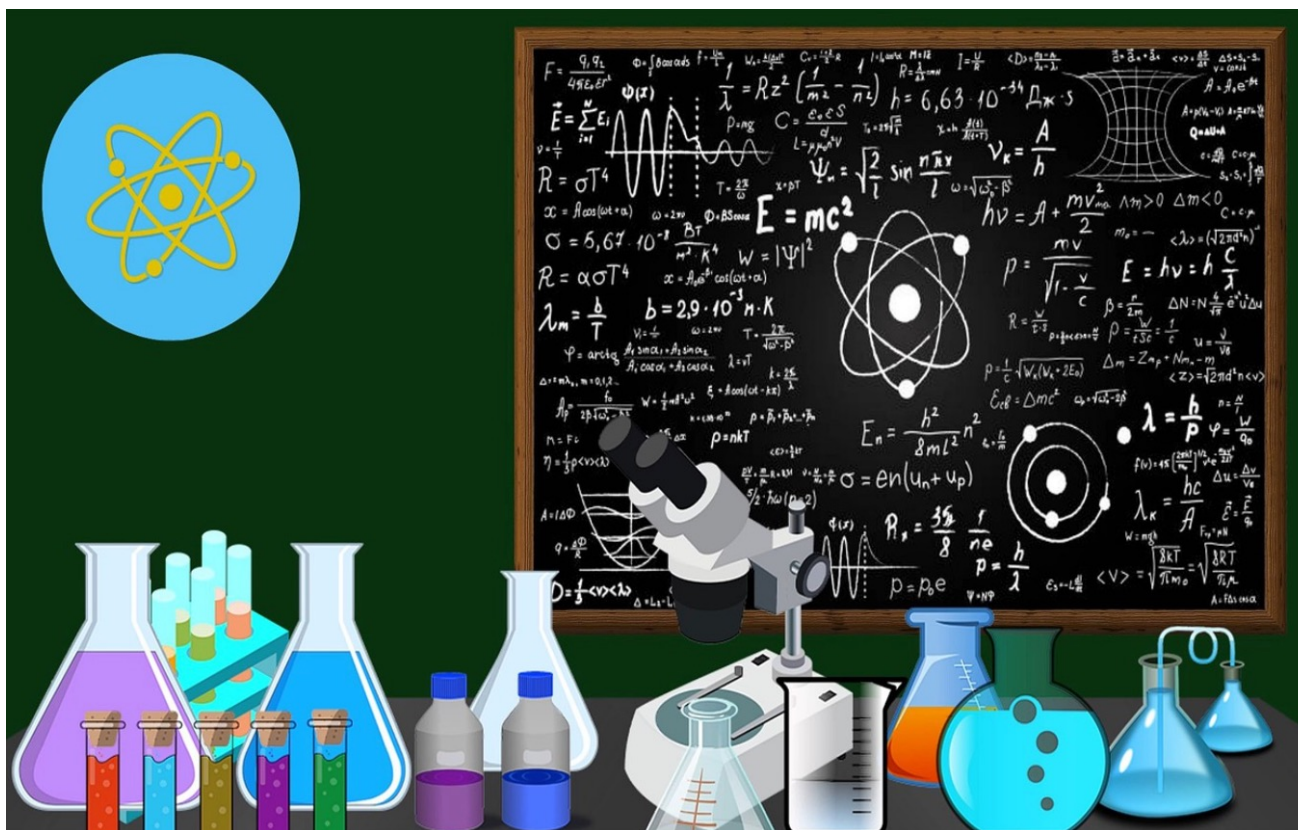
- Some slides appear blank because they have been removed
- Student versions have portions of the text removed which is given in the teacher version and appear as _____ which the students are to fill in
- Other slides may have on them, this represents writing that has been removed on this sample

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What is Stoichiometry?

Stoichiometry deals with calculations involving chemical reactions.



Chemical Reactions and the Mole

- Using the mole and balanced chemical equations, mass relationships can be determined for a chemical reaction

- _____



Steps to Stoichiometry

1. _____

2. _____

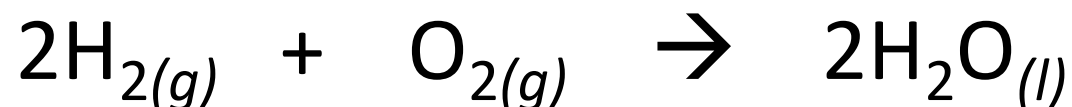
3. _____

4. _____

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Mole Ratio

Look at the creation of water:



The ratio of the components in this reaction is

2 molecules of $\text{H}_{2(g)}$: 1 molecule of $\text{O}_{2(g)}$: 2 molecules of $\text{H}_2\text{O}_{(l)}$

Mole Ratio

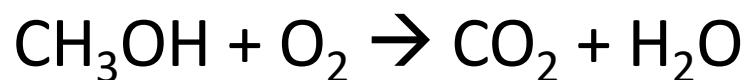


- _____
- Recall that Avogadro's Constant (6.02×10^{23}) is just a number. The mole represents the amount of an element or compound.
- When looking at the production of water we can use the same ratio



Check Your Understanding

Methanol burns in air according to the equation;



If 209 g of methanol are used up in the combustion, what mass of water is produced?

1. Write out the proper balanced equation
2. Determine moles of known substance
3. Determine moles of unknown substance from ratio
4. Determine quantity (mass or number of particles) using mole theory

A chart might help you to solve these questions

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Check Your Understanding

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