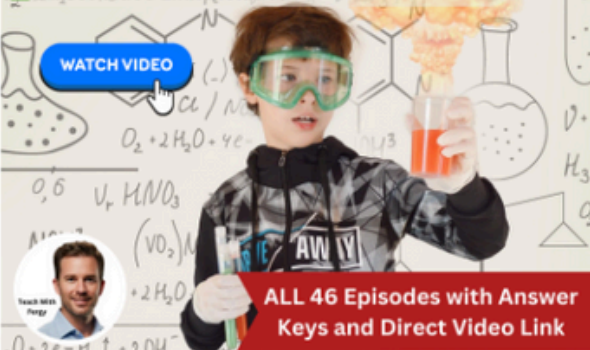


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Crash Course Chemistry #3: The Creation of Chemistry - The Fundamental Laws

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1. The _____ theory suggested that substances contained an invisible element released during burning.
2. Antoine Lavoisier helped define the _____ system to standardize measurements in chemistry.
3. The Law of Conservation of Mass used to be known as _____'s Law.
4. Lavoisier discovered that both ash and diamond contained pure _____.
5. The gas Lavoisier named _____ was shown to be vital for combustion.
6. Lavoisier proved that combustion and respiration both consumed _____ and produced carbon dioxide.
7. The transformation from alchemy to chemistry began with the realization that _____ is conserved during chemical reactions.
8. Joseph Proust established the _____, stating that chemical compounds contain elements in consistent ratios.
9. John Dalton used data to support the existence of _____ as indivisible units of matter.
10. Dalton initially believed that atoms and _____ were the same thing.
11. Gay-Lussac's experiments indicated that oxygen could split into _____ atoms.
12. Avogadro proposed that equal volumes of gases at the same temperature and pressure contain the same number of _____.
13. Avogadro's work helped determine the relative _____ of atoms and molecules.
14. The proposal that a gas like water vapor comes from the combination of two hydrogen atoms and one _____ atom challenged Dalton's model.
15. The number of molecules in a mole, later known as _____'s number, is crucial in chemistry.

Crash Course Chemistry #3: The Creation of Chemistry - The Fundamental Laws KEY

1. The **phlogiston** theory suggested that substances contained an invisible element released during burning.
2. Antoine Lavoisier helped define the **metric** system to standardize measurements in chemistry.
3. The Law of Conservation of Mass used to be known as **Lavoisier's** Law.
4. Lavoisier discovered that both ash and diamond contained pure **carbon**.
5. The gas Lavoisier named **hydrogen** was shown to be vital for combustion.
6. Lavoisier proved that combustion and respiration both consumed **oxygen** and produced carbon dioxide.
7. The transformation from alchemy to chemistry began with the realization that **mass** is conserved during chemical reactions.
8. Joseph Proust established the **Law of Definite Proportions**, stating that chemical compounds contain elements in consistent ratios.
9. John Dalton used data to support the existence of **atoms** as indivisible units of matter.
10. Dalton initially believed that atoms and **molecules** were the same thing.
11. Gay-Lussac's experiments indicated that oxygen could split into **two** atoms.
12. Avogadro proposed that equal volumes of gases at the same temperature and pressure contain the same number of **molecules**.
13. Avogadro's work helped determine the relative **mass** of atoms and molecules.
14. The proposal that a gas like water vapor comes from the combination of two hydrogen atoms and one **oxygen** atom challenged Dalton's model.
15. The number of molecules in a mole, later known as **Avogadro's** number, is crucial in chemistry.

15 Questions

Answer Key Included

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