

# UPGRADE to the Bundle

## Crash Course - CHEMISTRY 46 Episode Series Bundle

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### Crash Course Chemistry #22: Types of Chemical Bonds

Name: \_\_\_\_\_ Date: \_\_\_\_\_

You can access the video here → <https://youtu.be/QXT40VM4vXI?si=DzBvJCr7m1gYK7ms>

- The ideal distance between two nuclei at the point of minimum energy is called the \_\_\_\_\_.
- A bond where electrons are shared equally between atoms is called a \_\_\_\_\_ covalent bond.
- A bond formed by the transfer of electrons from one atom to another is called an \_\_\_\_\_ bond.
- The strength with which an atom holds shared electrons is known as \_\_\_\_\_.
- When electrons are more attracted to one side of a bond, creating partial charges, the bond is said to have \_\_\_\_\_.
- The formula used to calculate the energy between two ions in an ionic bond is called \_\_\_\_\_ law.
- Bonds form because atoms seek to minimize their overall \_\_\_\_\_.
- In an \_\_\_\_\_ bond, electrons are transferred from a metal to a non-metal.
- Covalent bonds typically form between \_\_\_\_\_ and sometimes \_\_\_\_\_.
- When the electrons in a bond are not shared equally, the bond is called a \_\_\_\_\_ covalent bond.
- \_\_\_\_\_ compounds are usually crystalline solids that dissolve in water and conduct electricity when dissolved.
- A \_\_\_\_\_ bond is formed when two identical atoms share electrons equally.
- The distance used in Coulomb's law calculations must be expressed in \_\_\_\_\_.
- In a covalent bond, electrons are \_\_\_\_\_ between atoms.
- \_\_\_\_\_ is the term used to describe the separation of charge within a bond or molecule.

### Crash Course Chemistry #22: Types of Chemical Bonds KEY

- The ideal distance between two nuclei at the point of minimum energy is called the **bond length**.
- A bond where electrons are shared equally between atoms is called a **non-polar** covalent bond.
- A bond formed by the transfer of electrons from one atom to another is called an **ionic** bond.
- The strength with which an atom holds shared electrons is known as **electronegativity**.
- When electrons are more attracted to one side of a bond, creating partial charges, the bond is said to have **polarity**.
- The formula used to calculate the energy between two ions in an ionic bond is called **Coulomb's law**.
- Bonds form because atoms seek to minimize their overall **energy**.
- In an **ionic** bond, electrons are transferred from a metal to a non-metal.
- Covalent bonds typically form between **non-metals** and sometimes **metalloids**.
- When the electrons in a bond are not shared equally, the bond is called a **polar** covalent bond.
- Ionic** compounds are usually crystalline solids that dissolve in water and conduct electricity when dissolved.
- A **non-polar** bond is formed when two identical atoms share electrons equally.
- The distance used in Coulomb's law calculations must be expressed in **nanometers**.
- In a covalent bond, electrons are **shared** between atoms.
- Polarity** is the term used to describe the separation of charge within a bond or molecule.

15 Questions

Answer Key Included

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