

# UPGRADE to the Bundle

## Crash Course - CHEMISTRY 46 Episode Series Bundle

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### Crash Course Chemistry #15: Partial Pressures & Vapor Pressure

Name: \_\_\_\_\_ Date: \_\_\_\_\_

You can access the video here → [https://youtu.be/\\_BqgCmYzA7si-fwAn1fPU6Lm2Oy5](https://youtu.be/_BqgCmYzA7si-fwAn1fPU6Lm2Oy5)

1. Dalton's discovery about gas mixtures is known as the Law of \_\_\_\_\_.
2. The ratio of the number of moles of a gas to the total moles in a mixture is called the \_\_\_\_\_.
3. The symbol for mole fraction is the lowercase Greek letter \_\_\_\_\_.
4. The formula for mole fraction is  $chi = n_i /$  \_\_\_\_\_.
5. The Ideal Gas Law is expressed as  $PV =$  \_\_\_\_\_.
6. Collecting a gas over water introduces \_\_\_\_\_ that must be subtracted to find the gas's pressure.
7. Vapor pressure of water increases with increasing \_\_\_\_\_.
8. The gas law demonstrated by summing individual gas pressures is called the Law of \_\_\_\_\_.
9. When gases are mixed, the total pressure equals the \_\_\_\_\_ of the partial pressures.
10. The reaction of acetic acid and baking soda produces \_\_\_\_\_ gas.
11. The molar mass of  $CO_2$  is \_\_\_\_\_ grams per mole.
12. Pressure created by water vapor is known as \_\_\_\_\_.
13. The collected  $CO_2$  volume in the experiment was approximately \_\_\_\_\_ milliliters.
14. Dalton's atomic theory built on the work of Joseph Louis \_\_\_\_\_.
15. The gas used in scuba tanks to help prevent decompression sickness is \_\_\_\_\_.

### Crash Course Chemistry #15: Partial Pressures & Vapor Pressure KEY

1. Dalton's discovery about gas mixtures is known as the Law of **Partial Pressures**.
2. The ratio of the number of moles of a gas to the total moles in a mixture is called the **mole fraction**.
3. The symbol for mole fraction is the lowercase Greek letter **chi**.
4. The formula for mole fraction is  $chi = n_i / n \text{ total}$ .
5. The Ideal Gas Law is expressed as  $PV = nRT$ .
6. Collecting a gas over water introduces **vapor pressure** that must be subtracted to find the gas's pressure.
7. Vapor pressure of water increases with increasing **temperature**.
8. The gas law demonstrated by summing individual gas pressures is called the Law of **Partial Pressures**.
9. When gases are mixed, the total pressure equals the **sum** of the partial pressures.
10. The reaction of acetic acid and baking soda produces **carbon dioxide** gas.
11. The molar mass of  $CO_2$  is **44** grams per mole.
12. Pressure created by water vapor is known as **vapor pressure**.
13. The collected  $CO_2$  volume in the experiment was approximately **90** milliliters.
14. Dalton's atomic theory built on the work of Joseph Louis **Proust**.
15. The gas used in scuba tanks to help prevent decompression sickness is **helium**.

15 Questions

Answer Key Included

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