

UPGRADE to the Bundle

Crash Course - CHEMISTRY 46 Episode Series Bundle

- ✓ Classroom Ready - Just Print and Go!
- ✓ Active Engagement - Keep Students' Attention
- ✓ Direct Video Link Included

WATCH VIDEO

ALL 46 Episodes with Answer
Keys and Direct Video Link

CLICK HERE



Crash Course Chemistry #36: Electrochemistry

Name: _____ Date: _____

You can access the video here → <https://youtu.be/TY4IUsholig?si=etmbMLZFconPdRC>

1. The type of chemistry that occurs in batteries is called _____.
2. Redox reactions involve the transfer of _____.
3. The flow of electrons through a conductor can be used to do _____.
4. The electrical potential of a redox reaction is also known as _____.
5. Redox reactions are typically broken down into _____ reactions.
6. In an alkaline battery, zinc is _____ and manganese is _____.
7. Alkaline batteries get their name from the presence of _____.
8. The terminal where electrons build up is the _____.
9. The terminal where electrons are drawn from is the _____.
10. Alkaline batteries are a type of _____ cell.
11. The component that completes the circuit in a galvanic cell by allowing ion flow is the _____.
12. The standard reduction potential is always written as if the substance is being _____.
13. The standard reduction potential for hydrogen is defined as _____ volts.
14. If the voltage of a redox reaction is positive, it will proceed _____ under standard conditions.
15. The opposite of a galvanic cell, which uses electricity to drive a reaction, is called an _____ cell.

Crash Course Chemistry #36: Electrochemistry KEY

1. The type of chemistry that occurs in batteries is called **electrochemistry**.
2. Redox reactions involve the transfer of **electrons**.
3. The flow of electrons through a conductor can be used to do **work**.
4. The electrical potential of a redox reaction is also known as **voltage**.
5. Redox reactions are typically broken down into **half reactions**.
6. In an alkaline battery, zinc is **oxidized** and manganese is **reduced**.
7. Alkaline batteries get their name from the presence of **potassium hydroxide**.
8. The terminal where electrons build up is the **cathode**.
9. The terminal where electrons are drawn from is the **anode**.
10. Alkaline batteries are a type of **galvanic** cell.
11. The component that completes the circuit in a galvanic cell by allowing ion flow is the **salt bridge**.
12. The standard reduction potential is always written as if the substance is being **reduced**.
13. The standard reduction potential for hydrogen is defined as **0.00** volts.
14. If the voltage of a redox reaction is positive, it will proceed **spontaneously** under standard conditions.
15. The opposite of a galvanic cell, which uses electricity to drive a reaction, is called an **electrolytic** cell.

15 Questions

Answer Key Included

Save 85% per worksheet!