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Crash Course Chemistry #29: Equilibrium Equations

Name: _____ Date: _____

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1. A reaction reaches _____ when the forward and reverse reactions occur at the same rate.
2. You can affect chemical equilibrium by changing concentration, temperature, or _____.
3. The _____ Process is used to synthesize ammonia by combining nitrogen and hydrogen.
4. The _____ constant is a number that represents the molar ratio of products to reactants at equilibrium.
5. In an equilibrium expression, coefficients from the balanced equation become _____.
6. Square brackets in a chemical formula represent _____, or moles per liter.
7. The value of the equilibrium constant is only valid at a specific _____.
8. A _____ table helps organize information to calculate equilibrium concentrations.
9. In a RICE table, the "C" row shows how much each substance's concentration will _____.
10. If a quadratic equation appears during a Keq calculation, you solve it using the _____ formula.
11. The concentration of hydrogen ions affects the equilibrium significantly if their coefficient is greater than one because their concentration is _____.
12. At equilibrium, the ratio of products to reactants raised to their coefficients must equal the _____ constant.
13. When a system at equilibrium is stressed, it shifts to _____ the stress.
14. The initial concentration of products is often zero because the reaction hasn't yet _____.
15. Changing the pressure in a reaction involving gases shifts the _____ of equilibrium.

Crash Course Chemistry #29: Equilibrium Equations KEY

1. A reaction reaches **equilibrium** when the forward and reverse reactions occur at the same rate.
2. You can affect chemical equilibrium by changing concentration, temperature, or **pressure**.
3. The **Haber** Process is used to synthesize ammonia by combining nitrogen and hydrogen.
4. The **equilibrium** constant is a number that represents the molar ratio of products to reactants at equilibrium.
5. In an equilibrium expression, coefficients from the balanced equation become **exponents**.
6. Square brackets in a chemical formula represent **molarity**, or moles per liter.
7. The value of the equilibrium constant is only valid at a specific **temperature**.
8. A **RICE** table helps organize information to calculate equilibrium concentrations.
9. In a RICE table, the "C" row shows how much each substance's concentration will **change**.
10. If a quadratic equation appears during a Keq calculation, you solve it using the **quadratic** formula.
11. The concentration of hydrogen ions affects the equilibrium significantly if their coefficient is greater than one because their concentration is **squared**.
12. At equilibrium, the ratio of products to reactants raised to their coefficients must equal the **equilibrium** constant.
13. When a system at equilibrium is stressed, it shifts to **counteract** the stress.
14. The initial concentration of products is often zero because the reaction hasn't yet **started**.
15. Changing the pressure in a reaction involving gases shifts the **position** of equilibrium.

15 Questions

Answer Key Included

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