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Crash Course Chemistry #28: Equilibrium

Name: _____ Date: _____

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1. In science, the word for balance is _____.
2. A reaction where products convert back into reactants is called a _____ reaction.
3. When the forward and reverse reactions occur at the same rate, the system is said to be at _____.
4. Chemical reactions that can proceed in both directions without extra energy are called _____ reactions.
5. A _____ arrow is used to represent reversible reactions in chemical equations.
6. Le Châtelier's _____ explains how systems at equilibrium respond to stress.
7. Adding more reactants to a system at equilibrium shifts the reaction to the _____.
8. Decreasing the pressure in a gas reaction shifts the equilibrium toward the side with more _____.
9. Increasing the pressure of a reaction involving gases favors the side with fewer _____.
10. _____ reactions are favored at higher temperatures because they absorb heat.
11. The Haber Process is used to produce _____ from nitrogen and hydrogen.
12. Removing a product from a system at equilibrium shifts the reaction toward the _____.
13. The concentration of substances, temperature, and _____ can all affect chemical equilibrium.
14. In the Haber Process, increasing the pressure shifts the reaction to the _____, producing more ammonia.
15. A system at equilibrium that is stressed will shift to _____ the stress.

Crash Course Chemistry #28: Equilibrium KEY

1. In science, the word for balance is **equilibrium**.
2. A reaction where products convert back into reactants is called a **reverse** reaction.
3. When the forward and reverse reactions occur at the same rate, the system is said to be at **chemical equilibrium**.
4. Chemical reactions that can proceed in both directions without extra energy are called **reversible** reactions.
5. A **double** arrow is used to represent reversible reactions in chemical equations.
6. Le Châtelier's **Principle** explains how systems at equilibrium respond to stress.
7. Adding more reactants to a system at equilibrium shifts the reaction to the **right**.
8. Decreasing the pressure in a gas reaction shifts the equilibrium toward the side with more **gas molecules**.
9. Increasing the pressure of a reaction involving gases favors the side with fewer **gas molecules**.
10. **Endothermic** reactions are favored at higher temperatures because they absorb heat.
11. The Haber Process is used to produce **ammonia** from nitrogen and hydrogen.
12. Removing a product from a system at equilibrium shifts the reaction toward the **right**.
13. The concentration of substances, temperature, and **pressure** can all affect chemical equilibrium.
14. In the Haber Process, increasing the pressure shifts the reaction to the **right**, producing more ammonia.
15. A system at equilibrium that is stressed will shift to **minimize** the stress.

15 Questions

Answer Key Included

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