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Crash Course Chemistry #9: Precipitation Reactions

Name: _____ Date: _____

You can access the video here → <https://youtu.be/1la16dy3Th7?si=MPGrDfL4MKQajSY>

1. Precipitation reactions occur when chemicals in a solution react to form a _____.
2. A key factor in precipitation reactions is a substance's _____ in water.
3. Ionic compounds consist of positively charged _____ and negatively charged _____.
4. _____ is the process of a solid forming out of a solution.
5. Silver chloride is an example of a(n) _____ compound that does not dissolve in water.
6. Sodium chloride, commonly known as _____, dissolves easily in water.
7. Precipitation reactions are used to extract _____ from mineral-rich solutions.
8. When writing chemical equations, (aq) stands for _____ and (s) stands for _____.
9. Compounds that do not change during a chemical reaction are called _____ ions.
10. A net ionic equation shows only the _____ participants in a reaction.
11. The Latin word for silver is _____, which is why its symbol is Ag.
12. Nitrate, sulfate, and phosphate ions are always _____.
13. To calculate how much reactant is needed, chemists use a _____ mass equation.
14. A _____ ounce equals 31.1 grams and is used to weigh precious metals.
15. A redox reaction, short for _____, is needed to isolate pure silver from silver chloride.

Crash Course Chemistry #9: Precipitation Reactions KEY

1. Precipitation reactions occur when chemicals in a solution react to form a **solid**.
2. A key factor in precipitation reactions is a substance's **solubility** in water.
3. Ionic compounds consist of positively charged **cations** and negatively charged **anions**.
4. **Precipitation** is the process of a solid forming out of a solution.
5. Silver chloride is an example of a(n) **insoluble** compound that does not dissolve in water.
6. Sodium chloride, commonly known as **salt**, dissolves easily in water.
7. Precipitation reactions are used to extract **silver** from mineral-rich solutions.
8. When writing chemical equations, (aq) stands for **aqueous** and (s) stands for **solid**.
9. Compounds that do not change during a chemical reaction are called **spectator** ions.
10. A net ionic equation shows only the **active** participants in a reaction.
11. The Latin word for silver is **argentum**, which is why its symbol is Ag.
12. Nitrate, sulfate, and phosphate ions are always **anions**.
13. To calculate how much reactant is needed, chemists use a **molar** mass equation.
14. A **troy** ounce equals 31.1 grams and is used to weigh precious metals.
15. A redox reaction, short for **reduction-oxidation**, is needed to isolate pure silver from silver chloride.

15 Questions

Answer Key Included

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