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Crash Course Chemistry #32: Kinetics

Name: _____ Date: _____

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1. Chemical reactions depend on _____ between molecules and atoms.
2. The minimum energy required for a reaction to occur is called _____ energy.
3. The field of chemistry that studies the rates of chemical reactions is called _____.
4. Molecules must collide with enough energy and the correct _____ for a reaction to occur.
5. The actual speed of the particles is their _____ energy.
6. The formula that shows the relationship between reactant concentrations and reaction rate is called the _____.
7. The constant in a rate law equation is usually represented by the letter _____.
8. In a rate law, the exponents must be determined _____.
9. The overall order of a reaction is found by _____ the exponents in the rate law.
10. A chemical that speeds up a reaction without being consumed is a _____.
11. A _____ is a reaction step that determines the speed of the entire reaction.
12. The slowest step in a multi-step reaction has the highest _____ energy.
13. Enzymes are biological examples of _____.
14. A reaction reaches _____ when its forward and reverse reaction rates are equal.
15. A collision that doesn't result in a reaction is considered _____.

Crash Course Chemistry #32: Kinetics KEY

1. Chemical reactions depend on **collisions** between molecules and atoms.
2. The minimum energy required for a reaction to occur is called **activation** energy.
3. The field of chemistry that studies the rates of chemical reactions is called **kinetics**.
4. Molecules must collide with enough energy and the correct **orientation** for a reaction to occur.
5. The actual speed of the particles is their **kinetic** energy.
6. The formula that shows the relationship between reactant concentrations and reaction rate is called the **rate law**.
7. The constant in a rate law equation is usually represented by the letter **k**.
8. In a rate law, the exponents must be determined **experimentally**.
9. The overall order of a reaction is found by **adding** the exponents in the rate law.
10. A chemical that speeds up a reaction without being consumed is a **catalyst**.
11. A **rate-determining step** is a reaction step that determines the speed of the entire reaction.
12. The slowest step in a multi-step reaction has the highest **activation** energy.
13. Enzymes are biological examples of **catalysts**.
14. A reaction reaches **equilibrium** when its forward and reverse reaction rates are equal.
15. A collision that doesn't result in a reaction is considered **ineffective**.

15 Questions

Answer Key Included

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