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Crash Course Chemistry #33: Doing Solids

Name: _____ Date: _____

You can access the video here → <https://youtu.be/bzr-bytSXIA?si=jHyOoAhmIqnn4dt>

1. Solids, like all matter, get their characteristics from the arrangement of their electrons, chemical bonds, and _____ forces.
2. There are two main classes of solids: _____ and amorphous.
3. The atoms and molecules in _____ solids are arranged in a random, disordered way.
4. The term _____ means "without shape," and refers to solids that lack a regular structure.
5. _____ solids melt gradually over a range of temperatures rather than at a specific point.
6. Crystalline solids tend to break along _____ lines due to their orderly arrangement.
7. Amorphous solids are _____, meaning they respond to stress the same way in every direction.
8. Crystalline solids are _____, meaning they respond differently depending on the direction of stress.
9. The three types of crystalline solids are molecular, _____, and atomic.
10. _____ solids are made of covalent compounds held together by weak Van der Waals forces.
11. _____ solids, like table salt, are made of charged ions and typically have high melting points.
12. Atomic solids include network solids, group 18 solids, and _____.
13. Metals are crystalline solids whose atoms are arranged in structures called _____ packing arrangements.
14. In metals, a sea of _____ allows for high conductivity and malleability.
15. The specific melting points of crystalline solids are due to the uniform _____ between their particles.

Crash Course Chemistry #33: Doing Solids KEY

1. Solids, like all matter, get their characteristics from the arrangement of their electrons, chemical bonds, and **intermolecular** forces.
2. There are two main classes of solids: **crystalline** and amorphous.
3. The atoms and molecules in **amorphous** solids are arranged in a random, disordered way.
4. The term **amorphous** means "without shape," and refers to solids that lack a regular structure.
5. **Amorphous** solids melt gradually over a range of temperatures rather than at a specific point.
6. Crystalline solids tend to break along **straight** lines due to their orderly arrangement.
7. Amorphous solids are **isotropic**, meaning they respond to stress the same way in every direction.
8. Crystalline solids are **anisotropic**, meaning they respond differently depending on the direction of stress.
9. The three types of crystalline solids are molecular, **ionic**, and atomic.
10. **Molecular** solids are made of covalent compounds held together by weak Van der Waals forces.
11. **Ionic** solids, like table salt, are made of charged ions and typically have high melting points.
12. Atomic solids include network solids, group 18 solids, and **metals**.
13. Metals are crystalline solids whose atoms are arranged in structures called **closest** packing arrangements.
14. In metals, a sea of **electrons** allows for high conductivity and malleability.
15. The specific melting points of crystalline solids are due to the uniform **bonds** between their particles.

15 Questions

Answer Key Included

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