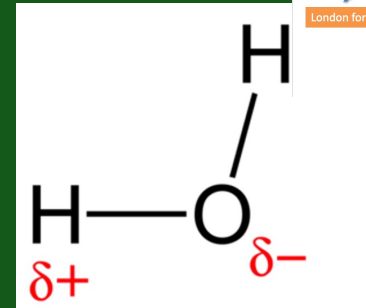
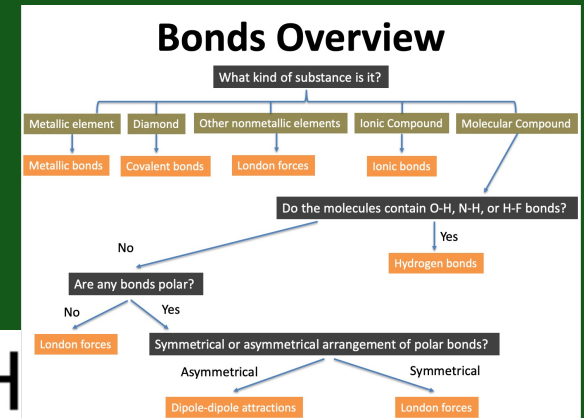
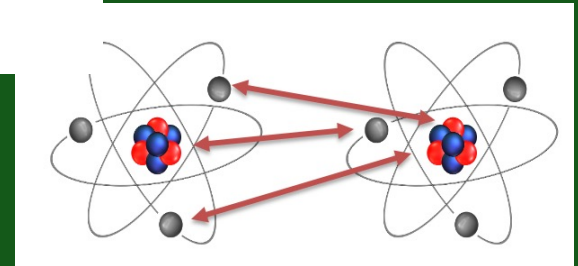
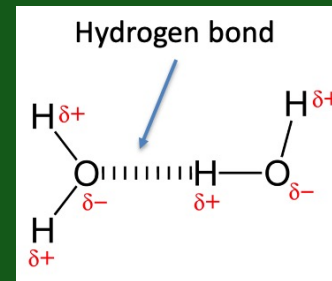


Intermolecular Forces

- The Four Types of Intermolecular Forces
- Dispersion Forces
- Molecular Dipole
- Dipole-Dipole Forces
- Impact of Intermolecular Forces on Melting and Boiling points
- Hydrogen Bonds
- Bonding Type Overview



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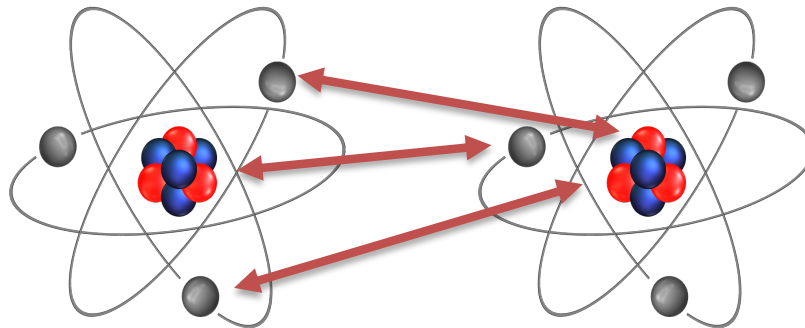


Intermolecular Forces

- Intermolecular Forces
- The strength and type of these forces depends on the structure of the molecules involved
- They are much weaker than a covalent or ionic bonds that hold molecules together
- Types of Intermolecular Forces:
 1. Dispersion Forces
 2. Dipole-dipole Forces
 3. Hydrogen Bonds
 4. Ion-dipole forces

Dispersion Forces

- An atomic nucleus not only attracts the electron in its own molecule but also those in neighbouring molecules
- Exists between ALL molecules
- Due to the attraction of the electrons of one molecule to the positive nuclei in the surrounding molecules
- _____

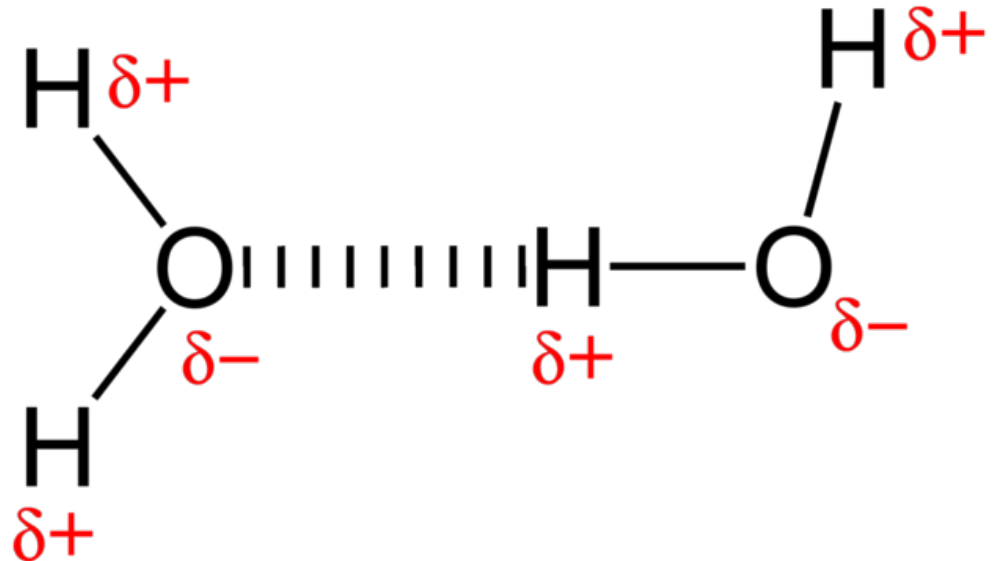


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Dipole-dipole Forces

- Attraction between dipoles due to simultaneous attraction between a dipole and its surrounding dipoles

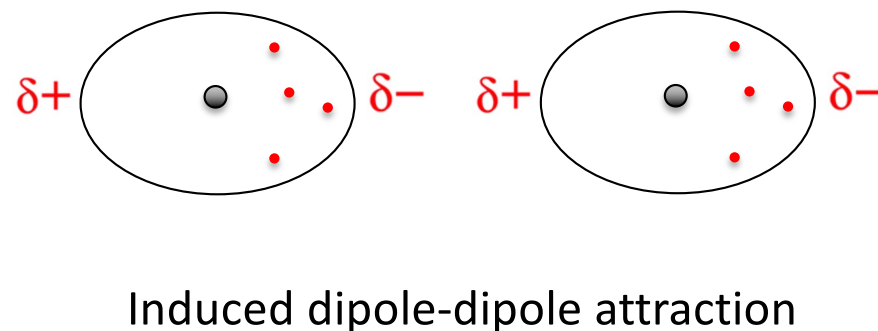
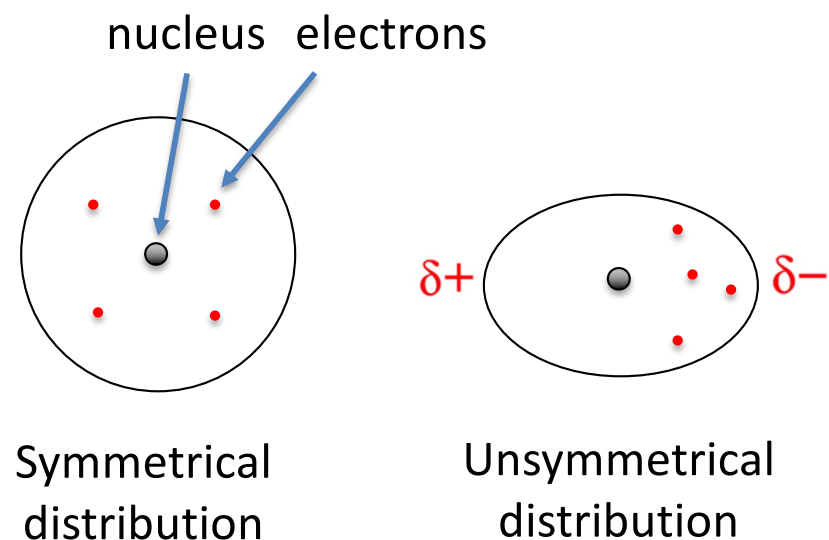
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London Forces

- The London dispersion force is the weakest intermolecular force
- It is a temporary attractive force that is created when adjacent electrons are positioned in a way that makes the atoms form temporary dipoles
- This is due to the electrons becoming dispersed unevenly about the nucleus
- This force is sometimes called an induced dipole-induced dipole attraction

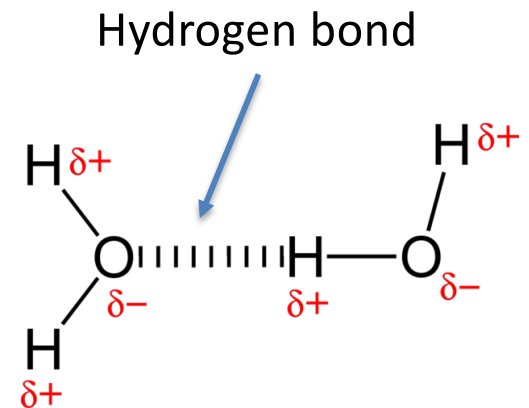
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Hydrogen Bonding

- _____
- The hydrogen in one of these molecules has a highly localized positive charge due to N, O and Fs high electronegativity
- The positive charge then becomes attracted to the localized negative charge of the N, O or F from another molecule
- It is a subcategory of dipole-dipole forces



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Bonds Overview

