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Crash Course Chemistry #12: The Ideal Gas Law

Name: _____ Date: _____

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- When a gas is compressed in a closed system, the pressure increases as the _____ decreases.
- The relationship between pressure and volume in gases is known as _____ Law.
- In the Ideal Gas Law, the letter "n" represents the number of _____.
- The Ideal Gas Law formula is written as _____.
- _____ discovered that volume divided by temperature equals a constant at constant pressure.
- Avogadro found that volume divided by number of _____ equals a constant at constant temperature and pressure.
- Pressure is commonly measured in _____ or atmospheres.
- The standard unit for pressure in the S.I. system is _____.
- The gas constant R has a value of _____ liters kPa/(mol K).
- Temperature in the Ideal Gas Law is a measure of the average _____ energy of particles.
- At STP, one mole of an ideal gas occupies _____ liters.
- STP stands for Standard _____ and _____.
- Absolute zero is defined as _____ kelvins or -273.15°C.
- When temperature decreases and gas condenses, pressure and _____ also decrease.
- Gases deviate from ideal behavior at low _____ or high _____.

Crash Course Chemistry #12: The Ideal Gas Law KEY

- When a gas is compressed in a closed system, the pressure increases as the **volume** decreases.
- The relationship between pressure and volume in gases is known as **Boyle's Law**.
- In the Ideal Gas Law, the letter "n" represents the number of **moles**.
- The Ideal Gas Law formula is written as **$PV = nRT$** .
- Charles** discovered that volume divided by temperature equals a constant at constant pressure.
- Avogadro found that volume divided by number of **moles** equals a constant at constant temperature and pressure.
- Pressure is commonly measured in **kilopascals** or atmospheres.
- The standard unit for pressure in the S.I. system is **pascals**.
- The gas constant R has a value of **8.3145** liters kPa/(mol K).
- Temperature in the Ideal Gas Law is a measure of the average **kinetic** energy of particles.
- At STP, one mole of an ideal gas occupies **22.4** liters.
- STP stands for Standard **Temperature and Pressure**.
- Absolute zero is defined as **0** kelvins or -273.15°C.
- When temperature decreases and gas condenses, pressure and **volume** also decrease.
- Gases deviate from ideal behavior at low **temperatures** or high **pressures**.

15 Questions

Answer Key Included

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