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Crash Course Chemistry #10: Redox Reactions

Name: _____ Date: _____

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1. The electron is to chemistry what _____ is to capitalism.
2. Electrons are responsible for allowing atoms to form _____.
3. Redox is a combination of the words _____ and _____.
4. In redox reactions, _____ is the gain of electrons.
5. The phrase "OILRIG" stands for "Oxidation is _____ of electrons, Reduction is _____ of electrons."
6. An atom's _____ number helps keep track of its electron ownership.
7. An element in its pure elemental form has an oxidation state of _____.
8. The oxidation state of a monatomic ion is equal to its _____.
9. Oxygen almost always has an oxidation state of _____, except in peroxides.
10. Hydrogen typically has an oxidation state of _____.
11. Fluorine usually has an oxidation state of _____, due to its high _____.
12. In a neutral compound, the sum of the oxidation states of all atoms must be _____.
13. In a polyatomic ion, the total oxidation state must equal the ion's _____.
14. The _____ process is a redox reaction used to produce ammonia.
15. Silver is _____ in the silver mirror reaction, forming a shiny metallic coating.

Crash Course Chemistry #10: Redox Reactions KEY

1. The electron is to chemistry what **money** is to capitalism.
2. Electrons are responsible for allowing atoms to form **molecules**.
3. Redox is a combination of the words **reduction** and **oxidation**.
4. In redox reactions, **reduction** is the gain of electrons.
5. The phrase "OILRIG" stands for "Oxidation is **Loss** of electrons, Reduction is **Gain** of electrons."
6. An atom's **oxidation** number helps keep track of its electron ownership.
7. An element in its pure elemental form has an oxidation state of **zero**.
8. The oxidation state of a monatomic ion is equal to its **charge**.
9. Oxygen almost always has an oxidation state of **-2**, except in peroxides.
10. Hydrogen typically has an oxidation state of **+1**.
11. Fluorine usually has an oxidation state of **-1**, due to its high **electronegativity**.
12. In a neutral compound, the sum of the oxidation states of all atoms must be **zero**.
13. In a polyatomic ion, the total oxidation state must equal the ion's **charge**.
14. The **Haber** process is a redox reaction used to produce ammonia.
15. Silver is **reduced** in the silver mirror reaction, forming a shiny metallic coating.

15 Questions

Answer Key Included

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